

1. Record Nr.	UNINA9911019420803321
Autore	Liu Shubin
Titolo	Exploring Chemical Concepts Through Theory and Computation
Pubbl/distr/stampa	Newark : , : John Wiley & Sons, Incorporated, , 2024 ©2024
ISBN	9783527843411 3527843418 9783527843435 3527843434 9783527843428 3527843426
Edizione	[1st ed.]
Descrizione fisica	1 online resource (594 pages)
Disciplina	542.8
Soggetti	Computational chemistry Quantum chemistry
Lingua di pubblicazione	Inglese
Formato	Materiale a stampa
Livello bibliografico	Monografia
Nota di contenuto	Cover -- Title Page -- Copyright -- Contents -- Preface -- Foreword -- 10 Questions About Exploring Chemical Concepts Through Theory and Computation -- Chapter 1 Chemical Concepts from Molecular Orbital Theory -- 1.1 Introduction -- 1.2 Molecular Orbital Theory -- 1.3 Canonical Molecular Orbitals -- 1.4 Frontier Molecular Orbital Theory -- 1.5 Localized Molecular Orbitals -- 1.5.1 Orthogonal Localized Molecular Orbitals -- 1.6 Regularized Nonorthogonal Localized Molecular Orbitals -- 1.7 Molecular Orbitals -- Acknowledgment -- References -- Chapter 2 Chemical Concepts from Ab Initio Valence Bond Theory -- 2.1 Introduction -- 2.2 Ab Initio Valence Bond Theory -- 2.2.1 Valence Bond SelfConsistent Field Method -- 2.2.2 Rumer Structures -- 2.2.3 Orbitals in VB Wave Function -- 2.2.4 VB Methods Involving Dynamic Correlation -- 2.3 Chemical Concepts in VB Theory -- 2.3.1 Resonance Theory -- 2.3.2 Conjugation, Hyperconjugation, and Aromaticity -- 2.3.3 ElectronPair Bonding in Valence Bond Theory -- 2.3.4 Diabatic States in Valence Bond Theory -- 2.4 A Brief Guide to Perform VB Calculations -- 2.4.1

Preparing XMVB Input Files -- 2.4.2 Reading XMVB Output Files -- 2.5
 Concluding Remarks -- References -- Chapter 3 Chemical Concepts
 from Conceptual Density Functional Theory -- 3.1 Introduction -- 3.2
 The Fundamentals: Density Functional Theory (DFT) and Kohn-Sham
 DFT -- 3.3 The First Derivatives: The Electronic Chemical Potential and
 the Electron Density -- 3.4 The Second Derivatives: Chemical Hardness,
 Fukui Function, Linear Response Function, and Related Quantities --
 3.4.1 Chemical Hardness and Softness -- 3.4.2 The Fukui Function and
 the Dual Descriptor -- 3.4.3 Local Softness and Hardness -- 3.4.4 The
 Linear Response Function, Softness and Hardness Kernels -- 3.5
 Perturbational Perspective of Chemical Reactivity -- 3.6 Conclusions --
 Acknowledgment.
 References -- Chapter 4 Chemical Concepts from DensityBased
 Approaches in Density Functional Theory -- 4.1 Introduction -- 4.2
 Four DensityBased Frameworks -- 4.2.1 OrbitalFree DFT (OFDFT) --
 4.2.2 Conceptual DFT (CDFT) -- 4.2.3 DensityAssociated Quantities
 (DAQs) -- 4.2.4 InformationTheoretic Approach (ITA) -- 4.3
 Applications of DensityBased Approaches -- 4.3.1 Molecular Isomeric
 and Conformational Stability -- 4.3.2 Bonding and Noncovalent
 Interactions -- 4.3.3 Cooperation and Frustration -- 4.3.4
 Homochirality and Principle of Chirality Hierarchy -- 4.3.5
 Electrophilicity and Nucleophilicity -- 4.3.6 Regioselectivity and
 Stereoselectivity -- 4.3.7 Brønsted-Lowry Acidity and Basicity -- 4.3.8
 Aromaticity and Antiaromaticity -- 4.3.9 Molecular Properties (Frontier
 Orbitals, HOMO/LUMO Gap, Oxidation States, Polarizability) -- 4.4
 Concluding Remarks -- Acknowledgments -- References -- Chapter 5
 Chemical Bonding -- 5.1 Introduction -- 5.2 The Physical Mechanism
 of the Chemical Bond -- 5.3 Bonding Models -- 5.4 Bond Length and
 Bond Strength -- 5.5 Dative and ElectronSharing Bonds -- 5.6 Polar
 Bonds -- 5.7 Atomic Partial Charges and Atomic Electronegativity --
 5.8 Chemical Bonding in MainGroup Compounds: N₂, CO, BF, LiF --
 5.9 Chemical Bonding of the Heavier MainGroup Atoms -- 5.10
 Chemical Bonding in Transition Metal Complexes: M(CO)_n (M & equals --
 Ni, Fe, Cr, Ti, Ca -- n & equals -- 4 - 8) -- 5.11
 Summary -- Acknowledgments -- References -- Chapter 6 Partial
 Charges -- 6.1 Concept of Partial Charge -- 6.1.1 What is Partial
 Charge? -- 6.1.2 Theoretical Significances and Practical Applications of
 Partial Charge -- 6.1.3 Limitations of Partial Charge -- 6.1.4 What Is a
 Good Method of Calculating Partial Charges? -- 6.1.5 Classification of
 Partial Charge Calculation Methods -- 6.2 Methods of Calculating
 Partial Charges.
 6.2.1 Partial Charges Based on Wavefunction -- 6.2.1.1 Mulliken
 Method -- 6.2.1.2 MMPA Methods -- 6.2.1.3 Löwdin Method -- 6.2.1.4
 NPA Method -- 6.2.2 Partial Charges Based on Real Space Partition of
 Electron Density -- 6.2.2.1 AIM Method -- 6.2.2.2 Voronoi and VDD
 Methods -- 6.2.2.3 Hirshfeld Method -- 6.2.2.4 HirshfeldI Method --
 6.2.3 Partial Charges Based on Fitting Electrostatic Potential -- 6.2.3.1
 Common ESP Fitting Methods -- 6.2.3.2 RESP and Relevant Methods --
 6.2.4 Partial Charges Based on Equalization of Electronegativity --
 6.2.5 Partial Charges Based on Other Ideas -- 6.3 Partial Charges of
 Typical Molecules -- 6.4 Computer Codes for Evaluating Partial
 Charges -- 6.5 Concluding Remarks -- References -- Chapter 7 Atoms
 in Molecules -- 7.1 Introduction -- 7.2 The Quantum Theory of Atoms
 in Molecules (QTAIM) -- 7.3 QTAIM Atoms as Open Quantum Systems
 -- 7.3.1 Sector Density Operators of Quantum Atoms in Molecules --
 7.3.2 RDMs of Atoms in Molecules -- 7.4 Interacting Quantum Atoms
 (IQA) -- References -- Chapter 8 Effective Oxidation States Analysis --
 8.1 The Concept of Oxidation State -- 8.2 Oxidation State is Not

Related to the Partial Charge -- 8.3 The Molecular Orbital Picture of the Ionic Approximation -- 8.4 SpinResolved Effective Fragment Orbitals and Effective Oxidation States (EOS) Analysis -- 8.5 EOS Analysis from Different AIM Schemes -- 8.6 Summary -- References -- Chapter 9 Aromaticity and Antiaromaticity -- 9.1 Definition of Aromaticity -- 9.2 Physical Foundation -- 9.3 Measures of Aromaticity -- 9.3.1 Geometric Descriptors of Aromaticity -- 9.3.2 Energetic Descriptors of Aromaticity -- 9.3.3 Electronic Descriptors of Aromaticity -- 9.3.4 Magnetic Descriptors of Aromaticity -- 9.4 Rules of Aromaticity -- 9.4.1 Rules for TwoDimensional Aromaticity -- 9.4.2 Rules for ThreeDimensional Aromaticity.

9.5 Metallabenzenes and Related Compounds as an Example -- References -- Chapter 10 Acidity and Basicity -- 10.1 Introduction -- 10.2 Definitions and Theories -- 10.2.1 Arrhenius Theory -- 10.2.2 Brønsted-Lowry Theory -- 10.2.3 Lewis Theory -- 10.2.4 Usanovich Definition -- 10.2.5 Lux-Flood Definition -- 10.2.6 Solvent System Definition -- 10.3 CDFTBased Reactivity Descriptors -- 10.4 CDFT Based Electronic Structure Principles -- 10.4.1 Equalization Principles -- 10.4.2 Hard-Soft Acid-Base (HSAB) Principle -- 10.4.3 Maximum Hardness (MHP), Minimum Polarizability (MPP), and Minimum Electrophilicity (MEP) Principles -- 10.5 Systemics of Lewis Acid-Base Reactions: Drago-Wayland Equation -- 10.6 Strengths of Acid and Bases -- 10.6.1 Ionic Product -- 10.6.2 pH Scale -- 10.6.3 Ionization Constants -- 10.6.4 Proton Affinity -- 10.6.5 Electronegativity -- 10.6.6 Hardness -- 10.6.7 Electrophilicity -- 10.7 Effect of External Perturbation -- 10.7.1 Steric Effects -- 10.7.2 Solvent Effects -- 10.7.3 Periodicity -- 10.7.4 Inductive Effect -- 10.7.5 Resonance Effect -- 10.8 CDFT and Acidity -- 10.9 CDFT and ITA -- 10.10 Are Strong Brønsted Acids Necessarily Strong Lewis Acids? -- 10.11 Summary -- Acknowledgment -- Conflict of Interest -- References -- Chapter 11 Sigma Hole Supported Interactions: Qualitative Features, Various Incarnations, and Disputations -- 11.1 Introduction -- 11.1.1 What's in a Name - The Sigma Hole Terminology and Concept -- 11.1.2 Donor-Acceptor Interaction Continuum -- 11.2 Many Incarnations and Roles of a Single Phenomenon -- 11.2.1 Hydrogen Bonding -- 11.2.2 Halogen Bonding and Sigma Holes on Group 17 Atoms -- 11.2.2.1 Common Origins -- 11.2.2.2 Cases of Halogen Bonding -- 11.2.2.3 The Sigma Hole and the Whole Story -- 11.2.3 Chalcogens -- 11.2.4 Pnictogens -- 11.2.5 Tetrels -- 11.2.6 Triels.

11.3 Related Interactions Elsewhere in the Main Group -- 11.3.1 Group 2 -- 11.3.2 Group 1 -- 11.3.3 Group 18 -- 11.4 Contested Interpretations -- 11.5 Conclusions -- Acknowledgment -- References -- Chapter 12 On the Generalization of Marcus Theory for TwoState Photophysical Processes -- 12.1 Introduction -- 12.2 The Golden Rule Rate Expression -- 12.2.1 The Marcus Theory: The Classical Treatment -- 12.2.2 The Marcus-Levich-Jortner Expression: A Quantum Expression for HighFrequency Modes -- 12.2.3 The Förster Theory: Separating Donor and Acceptor Parts in FCWD -- 12.3 Application -- 12.3.1 Electron Transfer -- 12.3.2 SET: Using Spectra for FCWD -- 12.3.3 TET and Other Energy Transfer Process with Spin Exchange -- 12.4 Conclusion -- Acknowledgments -- References -- Chapter 13 Computational Modeling of CO₂ Reduction and Conversion via Heterogeneous and Homogeneous Catalysis -- 13.1 Introduction -- 13.2 Computational Methods -- 13.3 Activation and Reduction of CO₂ -- 13.3.1 Computational Catalyst Design -- 13.3.1.1 Doping of Metal and Nonmetal Atoms -- 13.3.1.2 Structural Modification -- 13.3.1.3 Application of an External Electric Field -- 13.3.2 Electrocatalytic Reduction of CO₂ -- 13.3.3 Hydrogenation Reduction of CO₂ -- 13.4

Catalytic Coupling of CO₂ with CH₄ -- 13.5 Homogeneous Catalytic Conversion of CO₂ -- 13.5.1 Catalytic CO₂ Fixation into Cyclic Carbonates -- 13.5.2 CO₂ Hydrogenation Catalyzed by Metal PNP Pincer Complexes -- 13.6 Conclusion and Outlook -- Acknowledgments -- References -- Chapter 14 Excited States in Conceptual DFT -- 14.1 Introduction -- 14.2 Exploring Ground State Properties Thanks to Excited States -- 14.2.1 Context and Justification -- 14.2.2 Chemical Hardness Revisited -- 14.2.3 StateSpecific Dual Descriptors -- 14.2.4 Polarization Interaction -- 14.3 Exploring the Reactivity of Excited States with Excited States. 14.3.1 Local Chemical Potential.

Sommario/riassunto

This comprehensive book, edited by Shubin Liu, delves into the exploration of chemical concepts through theoretical and computational chemistry. It covers a broad range of topics including molecular orbital theory, valence bond theory, density functional theory, and energy decomposition analysis. The book aims to provide insights into chemical bonding, partial charges, aromaticity, and various other fundamental chemical concepts. The chapters are contributed by numerous experts, offering both static and dynamic properties of chemical entities. This resource is invaluable for researchers, academics, and advanced students in the field of chemistry who are interested in understanding the theoretical underpinnings and computational approaches that define chemical principles and empirical laws.
