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Nota di contenuto	Table of Contents; Title; Copyright; Preface; Notations and Symbols; 1: Physico-Chemical Transformations and Equilibria; 1.1. Characteristic parameters of physico-chemical transformations; 1.2. Entropy production during the course of a transformation in a closed system; 1.3. Affinity of a transformation; 1.4. De Donder's inequality - direction of the transformations and equilibrium conditions; 1.5. Heats of transformation; 1.6. Set of points representing the equilibrium states of a transformation; 1.7. Closed systems accommodating multiple reactions 1.8. Direction of evolution and equilibrium conditions in an open system 1.9. Azeotropic transformations; 2: Properties of States of Physico-Chemical Equilibrium; 2.1. Laws of displacement of an equilibrium; 2.2. Properties of all the equilibria in a system; 2.3. Phase laws; 2.4. Indifferent states; 2.5. Thermodynamically-equivalent systems; 2.6. Stability of equilibria; 3: Molecular Chemical Equilibria; 3.1. Law of mass action - equilibrium constants; 3.2. Graphical representations of equilibria - pole diagrams; 3.3. Representation of the evolution of an equilibrium with the temperature 3.4. Binary diagrams for chemical equilibrium 3.5. Ternary diagrams of

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4.7. Determination of the equilibrium constants on the basis of spectral data and statistical thermodynamics  
4.8. Thermodynamic tables and databanks; 4.9. Estimation of thermodynamic data; 4.10. Thermodynamic calculations for complex systems; APPENDICES; Appendix 1: Recap on the Reference States of Solutions; A1.1. Concentration and molar fraction; A1.2. Chemical potentials and activity coefficients; A1.3. Characterization of the imperfection of a real solution by the excess Gibbs energy; Appendix 2: Recap of statistical thermodynamics; A2.1. The three branches of statistics  
A2.2. Partition functions of a molecule object  
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