1. Record Nr. UNINA9910536700903321

Autore	Elliot Patricia <1955->
Titolo	Debates in transgender, queer, and feminist theory : contested sites / / by Patricia Elliot
Pubbl/distr/stampa	Farnham, Surrey, UK ; ; Burlington, VT, : Ashgate Pub., c2010
ISBN	9786612892233 9781315576008 1315576007 9781317154334 1317154339 9781317154327 1317154320 9781282892231 1282892231 9781409403944 1409403947
Edizione	[1 ed.]
Descrizione fisica	1 online resource (205 p.)
Collana	Queer interventions
Classificazione	SOC032000
Disciplina	306.76/8
Soggetti	Gender nonconformity Feminist theory Gender identity Queer theory
Lingua di pubblicazione	Inglese
Formato	Materiale a stampa
Livello bibliografico	Monografia
Note generali	Description based upon print version of record.
Nota di bibliografia	Includes bibliographical references and index.
Nota di contenuto	Cover; Contents; Series Editor's Preface Trans-positions, Fugitive Poetics and Educated Hope; Preface and Acknowledgements; About the Author; Introduction Exploring Rifts in Transgender, Queer, and Feminist Theories; 1 Feminist Embattlement on the Field of Trans; 2 Revaluing Gender Diversity Beyond the Ts/Tg Hierarchy; 3 Desire and the "(Un)Becoming Other": The Question of Intelligibility; 4 Risking the Unfamiliar: Psychic Complexity in Theories of Transsexual Embodiment; 5 Still Not in Our Genes: Theorizing Complex Bodies; Conclusion Fielding Contested Desires; Bibliography; Index

Sommario/riassunto	Intersecting the domains of women's studies, sexuality, gender and transgender studies, Debates in Transgender, Queer, and Feminist Theory provides a critical analysis of key texts and theories, engaging in a dialogue with prominent theorists of transgendered identity, embodiment and sexual politics, and intervening in various aspects of a conceptually and politically difficult terrain.

2. Record Nr. UNINA9911020379503321

Autore	Rauk Arvi <1942->
Titolo	Orbital interaction theory of organic chemistry / / by Arvi Rauk
Pubbl/distr/stampa	New York, : Wiley-Interscience, 2001
ISBN	9786610264704 9781280264702 1280264705 9780470351062 0470351063 9780471461845 0471461849 9780471220411 0471220418
Edizione	[2nd ed.]
Descrizione fisica	1 online resource (360 p.)
Disciplina	547/.128
Soggetti	Molecular orbitals Physical organic chemistry
Lingua di pubblicazione	Inglese
Formato	Materiale a stampa
Livello bibliografico	Monografia
Note generali	"A Wiley-Interscience publication."
Nota di bibliografia	Includes bibliographical references (p. 313-324) and index.
Nota di contenuto	CONTENTS; PREFACE; 1 SYMMETRY AND STEREOCHEMISTRY; Purpose; Definition of a Group; Molecular Point Groups; Schoenflies Notation; Interrelations of Symmetry Elements; Type Classification; Isomerism and Measurements; Stereoisomerism of Molecules; Stereotopic Relationships of Groups in Molecules; Asymmetric Synthesis and Stereochemistry; NMR and Stereochemistry; Symmetry and Structural

Parameters; Note on Hybridization; Symmetry and Orbitals; Atomic Orbitals; Molecular and Group Orbitals; In What Combination?; 2 MOLECULAR ORBITAL THEORY; Introduction; Electronic Schrodinger Equation (A.1)
Fock Equations (A.42)The Basis Set (STO-3G, 6-31G*, and All That); Orbital Energies and Orbitals; Representation of MOs; Total Energies and the Hartree-Fock Limit; Successes and Failures of Hartree-Fock Theory; Beyond Hartree-Fock; Density Functional Theory; Geometry Optimization; Normal Coordinates and Harmonic Frequency Analysis; Zero Point Vibrational Energies; 3 ORBITAL INTERACTION THEORY; Relationship to Hartree-Fock Equations; Huckel Approximation; Orbital Energies and Total Electronic Energy; Case Study of a Two-Orbital Interaction; Case 1: [sub(A)] = [sub(B)], S[sub(AB)] = 0
Case 2: [sub(A)] = [sub(B)], [sub(AB)] > 0, [sub(AB)] « 1Case 3: [sub (A)] > [sub(B)], S[sub(AB)] = 0; Case 4: [sub(A)] > [sub(B)], S[sup(AB)] > 0; Effect of Overlap; Energetic Effect of Overlap; Orbital Effect of Overlap; First Look at Bonding; Relationship to Perturbation Theory; Generalizations for Intermolecular Interactions; Energy and Charge Distribution Changes from Orbital Interaction; Four-Electron, Two-Orbital Interaction; Three-Electron, Two-Orbital Interaction; Two-Electron, Two-Orbital Interaction; One-Electron, Two-Orbital Interaction; Zero-Electron, Two-Orbital Interaction
Interactions between Molecules: Many Electrons, Many OrbitalsGeneral Principles Governing the Magnitude of h[sub(AB)] and S[sub(AB)]; Interactions of MOs; Electrostatic Effects; Group Orbitals; Zero-Coordinated Atoms; Monocoordinated Atoms; Dicoordinated Atoms; Tricoordinated Atoms; Tetracoordinated Atoms; Assumptions for Application of Qualitative MO Theory; Example: Carbonyl Group; Construction of Interaction Diagram; Interpretation of Interaction Diagram; Chemical Reactivity; Why Does It Work and When Might it Not?; 4 SIGMA BONDS AND ORBITAL INTERACTION THEORY
C-X Bonds: X = C, N, O, F and X = F, Cl, Br, I Bonds: Homolytic versus Heterolytic Cleavage; Heterolytic Cleavage of Bonds Involving C or H; Homolytic Cleavage of Bonds Involving C or H; Homonuclear Bonds C-C, N-N, O-O, F-F, Cl-Cl, Br-Br, and I-I; Interactions of Bonds; Bonds as Electron Donors or Acceptors; Bonds as Electron Acceptors; As a Acceptor; As a Acceptor; Bonds as Electron Donors; As a Donor; As a Donor; Bonding in Cyclopropane; 5 SIMPLE HUCKEL MOLECULAR ORBITAL THEORY; Simple Huckel Assumptions
Charge and Bond Order in SHMO Theory: (S[sub(AB)] = 0, One Orbital per Atom)

Sommario/riassunto

A practical introduction to orbital interaction theory and its applications in modern organic chemistry Orbital interaction theory is a conceptual construct that lies at the very heart of modern organic chemistry. Comprising a comprehensive set of principles for explaining chemical reactivity, orbital interaction theory originates in a rigorous theory of electronic structure that also provides the basis for the powerful computational models and techniques with which chemists seek to describe and exploit the structures and thermodynamic and kinetic stabilities of molecules. Orbital Interaction