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Altri autori (Persone)	ZemaitisJoseph F. <b. 1940.>
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Nota di contenuto	Handbook of Aqueous Electrolyte Thermodynamics Theory & Application; TABLE OF CONTENTS; I INTRODUCTION; II THERMODYNAMICS OF SOLUTIONS; Basic Thermodynamic Functions; Solutions - Basic Definitions and Concepts; Equilibrium - Necessary Conditions; Activities, Activity Coefficients and Standard States; III EQUILIBRIUM CONSTANTS; Ionic and/or Reaction Equilibrium in Aqueous Solutions; Solubility Equilibria Between Crystals and Saturated Solutions; Vapor-Liquid Equilibria in Aqueous Solutions; Temperature Effects on the Equilibrium Constant Estimating Temperature Effects on Heat Capacity and Other Thermodynamic PropertiesEquilibrium Constants from Tabulated Data; Pressure Effects on the Equilibrium Constant; Appendix 3.1 - Criss and Cobble Parameters; IV ACTIVITY COEFFICIENTS OF SINGLE STRONG ELECTROLYTES; History; Limitations and Improvements to the Debye-Huckel Limiting Law; Further Refinements; Bromley's Method; Meissner's Method; Pitzer's Method; Chen's Method; Short Range Interaction Model; Long Range Interaction Model; Temperature Effects;

Bromley's Method; Meissner's Method; Pitzer's Method; Chen's Method; Application

Bromley's Method; Meissner's Method; Pitzer's Method; Chen's Method; NBS Smoothed Experimental Data; Test Cases; HCl; KCl; KOH; NaCl; NaOH; CaCl<sub>2</sub>; Na<sub>2</sub>SO<sub>4</sub>; MgSO<sub>4</sub>; Bromley's Extended Equation; MgSO<sub>4</sub> Test Case; Comparison of Temperature Effect Methods; Bromley; Meissner; Pitzer and Chen; Experimental Data; Test Cases; HCl at 50° Celsius; KCl at 80° Celsius; KOH at 80° Celsius; NaCl at 100 and 300° Celsius; NaOH at 35° Celsius; CaCl<sub>2</sub> at 108.85 and 201.85° Celsius; Na<sub>2</sub>SO<sub>4</sub> at 80° Celsius; MgSO<sub>4</sub> at 80° Celsius; Appendix 4.1 - Values for Guggenheim's Parameter

Table 1: Values for Uni-univalent Electrolytes; Table 2:  $\alpha$  and  $\beta$  Values of Bi-univalent and Uni-bivalent Electrolytes from Freezing Points;

Methods for Calculating; Appendix 4.2 - Bromley Interaction Parameters; Table 1: B Values at 25°C Determined by the Method of Least Squares on  $\log \gamma$  to  $I=6.0$  (or less if limited data); Table 2:

Individual Ion Values of B and  $\beta$  in Aqueous Solutions at 25°C; Table 3: Bivalent Metal Sulfates at 25°C; Appendix 4.3 - Meissner Parameters; Table: Average Values of Parameter  $q$  in Equation (4.46) for Selected Electrolytes; Appendix 4.4 - Pitzer Parameters

Table 1: Inorganic Acids, Bases and Salts of 1-1 Type; Table 2: Salts of Carboxylic Acids (1-1 Type); Table 3: Tetraalkylammonium Halides; Table 4: Sulfonic Acids and Salts (1-1 Type); Table 5: Additional 1-1 Type Organic Salts; Table 6: Inorganic Compounds of 2-1 Type; Table 7: Organic Electrolytes of 2-1 Type; Table 8: 3-1 Electrolytes; Table 9: 4-1 Electrolytes; Table 10: 5-1 Electrolytes; Table 11: 2-2 Electrolytes; Appendix 4.5 - Pitzer Parameter Derivatives; Table 1: Temperature Derivatives of Parameters for 1-1 Electrolytes Evaluated from Calorimetric Data

Table 2: Temperature Derivatives of Parameters for 2-1 and 1-2 Electrolytes Evaluated from Calorimetric Data

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## Sommario/riassunto

Expertise in electrolyte systems has become increasingly important in traditional CPI operations, as well as in oil/gas exploration and production. This book is the source for predicting electrolyte systems behavior, an indispensable "do-it-yourself" guide, with a blueprint for formulating predictive mathematical electrolyte models, recommended tabular values to use in these models, and annotated bibliographies. The final chapter is a general recipe for formulating complete predictive models for electrolytes, along with a series of worked illustrative examples. It can serve as a useful resea

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