1. Record Nr. UNISA996393536003316

Titolo	By the King. A proclamation for making currant certaine French coyne [[electronic resource]]
Pubbl/distr/stampa	Printed at Oxford, : by I. L[ichfield]. and W. T[urner]. for Bonham Norton, and Iohn Bill [in London], printers to the Kings most Excellent Maiestie, 1625
Descrizione fisica	1 sheet ([1] p.)
Altri autori (Persone)	Charles, King of England, <1600-1649.>
Soggetti	Coinage - Law and legislation - England Coins, French - Law and legislation
Lingua di pubblicazione	Inglese
Formato	Materiale a stampa
Livello bibliografico	Monografia
Note generali	Dated at end: Tichfeld the fourth day of September, in the first yeare of his Maiesties raigne ... . Steele notation: deare London, the Reproduction of original in the Bodleian Library, Oxford, England.
Sommario/riassunto	eebo-0014

2. Record Nr. UNINA9910140600803321

Autore	Ghali Edward
Titolo	Corrosion resistance of aluminum and magnesium alloys [[electronic resource]] : understanding, performance, and testing / / Edward Ghali
Pubbl/distr/stampa	Hoboken, N.J., : Wiley, c2010
ISBN	1-282-68420-5 9786612684203 0-470-53177-0 0-470-53176-2
Descrizione fisica	1 online resource (743 p.)
Collana	Wiley series on corrosion
Disciplina	620.1/8623
Soggetti	Aluminum alloys - Corrosion Magnesium alloys - Corrosion Corrosion and anti-corrosives
Lingua di pubblicazione	Inglese
Formato	Materiale a stampa
Livello bibliografico	Monografia
Note generali	Description based upon print version of record.
Nota di bibliografia	Includes bibliographical references and index.
Nota di contenuto	Corrosion Resistance of Aluminum and Magnesium Alloys; Contents; Preface; Acknowledgments; Part One Electrochemical Fundamentals and Active-Passive Corrosion Behaviors; 1. Fundamentals of Electrochemical Corrosion; Overview; A. Thermodynamic Considerations of Corrosion; 1.1. Electrolytic Conductance; 1.1.1. Faraday Laws; 1.2. Tendency to Corrosion; 1.3. The Electrochemical Interface; 1.3.1. Electric Double Layer; 1.3.2. Equivalent Circuit of the Electric Double Layer; 1.4. Nernst Equation; 1.5. Standard Potentials of Electrodes; 1.5.1. Standard States in Solution; 1.5.2. Hydrogen Electrode 1.5.3. Positive and Negative Signs of Potentials1.5.4. Graphical Presentation; B. Activity and Conductance of the Electrolyte; 1.6. Activity of the Electrolyte; 1.6.1. Constant and Degree of Dissociation; 1.6.2. Activity and Concentration; 1.6.3. Theory of More Concentrated Solutions; 1.6.4. Electrolytic Conduction; 1.7. Mobility of Ions; 1.7.1. Law of Additivity of Kohlrausch; 1.7.2. Ion Transport Number or Index; 1.8. Conductance; 1.9. Potential of Decomposition; C. The Different Types of Electrodes; 1.10. Gas Electrodes; 1.11. Metal-Metal Ion Electrodes; 1.11.1. Alloyed Electrodes

1.12. Metal-Insoluble Salt or Oxide Electrodes1.12.1. Metal-Insoluble Salt Electrodes; 1.12.2. Metal-Insoluble Oxide Electrodes; 1.13. Electrodes of Oxidation-Reduction; 1.14. Selective Ion Electrodes; 1.14.1. Glass Electrodes; 1.14.2. Copper Ion-Selective Electrodes; D. Electrochemical and Corrosion Cells; 1.15. Chemical Cells; 1.15.1. Chemical Cell with Transport; 1.15.2. Chemical Cell Without Transport; 1.16. Concentration Cells; 1.16.1. Concentration Cell with Difference of Activity at the Electrode and Electrolyte; 1.16.2. Junction Potential; 1.17. Solvent Corrosion Cells
1.17.1. Cathodic Oxidoreduction Reaction1.17.2. Displacement Cell; 1.17.3. Complexing Agent Cells; 1.17.4. Stray Current Corrosion Cell; 1.18. Temperature Differential Cells; 1.19. Overlapping of Different Corrosion Cells; E. Chemical and Electrochemical Corrosion; 1.20. Definition and Description of Corrosion; 1.21. Electrochemical and Chemical Reactions; 1.21.1. Electrochemical Corrosion; 1.21.2. Film-Free Chemical Interactions; References; 2. Aqueous and High-Temperature Corrosion; Overview; 2.1. Atmospheric Media; 2.1.1. Description; 2.1.2. Types of Corrosion
2.1.3. Atmospheric Contaminants2.1.4. Corrosion Prevention and Protection; 2.2. Aqueous Environments; 2.3. Organic Solvent Properties; 2.4. Underground Media; 2.5. Water Media Properties; 2.5.1. Water Composition; 2.5.2. The Oxidizing Power of Solution; 2.5.3. Scale Formation and Water Indexes; 2.6. Corrosion at High Temperatures; 2.6.1. Description; 2.6.2. The Pilling-Bedworth Ratio (PBR); 2.6.3. Kinetics of Formation; 2.6.4. Corrosion Behaviors of Some Alloys at Elevated Temperatures; References; 3. Active and Passive Behaviors of Aluminum and Magnesium and Their Alloys; Overview
3.1. Potential-pH Diagrams of Aluminum and Magnesium

Sommario/riassunto

Valuable information on corrosion fundamentals and applications of aluminum and magnesium Aluminum and magnesium alloys are receiving increased attention due to their light weight, abundance, and resistance to corrosion. In particular, when used in automobile manufacturing, these alloys promise reduced car weights, lower fuel consumption, and resulting environmental benefits. Meeting the need for a single source on this subject, Corrosion Resistance of Aluminum and Magnesium Alloys gives scientists, engineers, and students a one-stop reference for understanding both the corrosion f